

Amendments to the Specification

The first paragraph on page 1 should be amended as follows:

--This application is a continuation-in-part of International application PCT/US03/03688 Filed February 7, 2003, a continuation-in-part of U.S. non-provisional application no. 10/277,234 filed October 22, 2002, now abandoned, and claims the benefit of U.S. provisional application no. 60/357,330 filed February 15, 2002.--.

The paragraph describing Example 2 on page 12 should be amended as follows:

-- Example 2: A semi-bright tin-lead deposit is obtained by adding 1.5g/l of lead methane sulfonate to the solution of [[claim]] Example 1 and plating at the same conditions.--.

The paragraph describing Example 3 on page 13 should be amended as follows:

--Example 3: The formulation of Example 1 was used to plate tin on 250 pieces of 8mm diameter flat washers in a 2.5" by [[4"barrel]] 4" barrel, 140 ml of 2.5 mm diameter conductive balls were used as the media. The load was plated at 5A, 6.5V for 15 minutes. At the end of the plating cycle, none of the flat washers were fused together.--.

The first paragraph beneath Example 4 on page 14 should be amended as follows:

--The drawing figure illustrates the effect of the ascorbic acid to tin concentration ratio on part agglomeration during electroplating. Ascorbic acid baths having a pH of 4.2, 4.5 and 5 and using different ratios of ascorbic acid to tin were used to plate tin on 250 pieces of 8mm diameter flat washers in a 2.5" by [[4"barrel]] 4" barrel, 140 ml of 2.5 mm diameter conductive balls were used as the media. The load was plated at 5A, 6.5V for 15 minutes. At the end of the plating cycle, the number of single washer and agglomerated washers were counted and summed. Therefore, a count of 250 represents that no washers were agglomerated. If all the washer were agglomerated into groups of two the count would be 125. As can be seen from the figure, lower ratios of ascorbic acid to tin result in less agglomeration. This effect is particularly pronounced at the lower pH values, as it is seen that higher complexer levels are required to maintain the tin in solution at higher pH values. Therefore, the higher pH value results are still consistent with the concept of limiting the amount of complexer in excess of that required to maintain the tin in solution.--.